

Rajiv Gandhi University of Health Sciences, Karnataka
I Year B.Pharm Degree Examination – May 2017

Time: Three Hours

Max. Marks: 70 Marks

PHARMACEUTICAL INORGANIC CHEMISTRY
(Revised Scheme 3)

Q.P. CODE: 2605

Your answers should be specific to the questions asked
Draw neat labeled diagrams wherever necessary

LONG ESSAYS (Answer any Two)

2 x 10 = 20 Marks

1. Describe the principle and reactions involved in the assay of a) Ferrous sulphate b) Copper sulphate.
2. Discuss with examples the various sources of impurities in Pharmaceutical substances.
3. Discuss the mechanism of maintaining the Physiological acid-base balance. Write a note on biological and biochemical role of sodium and bicarbonate ions in the body.

SHORT ESSAYS (Answer any Six)

6 x 5 = 30 Marks

4. What are Antidotes? Give its classification with examples.
5. What are primary standards, give examples. What are the properties of an ideal primary standard?
6. Write the preparation and standardization of 0.1N HClO₄.
7. Explain the different sources of errors and the various methods to minimize them.
8. Discuss the role of fluoride in dental caries.
9. Give the pharmaceutical applications of non-aqueous titrations.
10. Write the procedure and principle with reactions for the limit test for Chloride.
11. Give the assay principle and medicinal uses of Boric acid.

SHORT ANSWERS

10 x 2 = 20 Marks

12. What are Acidifiers? Give example.
13. Role of Acetic acid and Ammonia in limit test for Heavy Metals.
14. Define Protective and Adsorbents.
15. ORS.
16. Define Titrant and titrate.
17. Define Pharmaceutical aid. Give example.
18. How will you prepare 250ml of 0.1 N NaOH solution?
19. Why dilute alcohol is used in the limit test for sulphate.
20. Write the method of preparation of magnesium sulphate.
21. Name some metal indicators.
